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Chapter 6 Solutions Engineering and Chemical Thermodynamics Wyatt Tenhaeff Milo Koretsky Department of Chemical Engineering Oregon State University [email protected] 2 6.1 (a) The Clausius-Clapeyron equation:  $d \ln P / dT = h_{vap} / (RT^2)$  or  $\ln P_2 / P_1 = -h_{vap} / R (1/T_2 - 1/T_1)$  so  $\ln P_2 = \ln P_1 - h_{vap} / R (1/T_2 - 1/T_1)$  so  $P_2 = P_1 \exp(-h_{vap} / R (1/T_2 - 1/T_1))$

### Chapter 6 Solutions - Chapter 6 Solutions Engineering and ...

Chapter 6 2 If the fluid is in thermodynamic equilibrium any thermodynamic variable for a pure substance, like pure water, can be written in terms of any two other thermodynamic variables, i.e.  $p = p(p, T)$  (6.1.1) where the functional relationship in depends on the substance.

### Chapter 6 Thermodynamics and the Equations of Motion

Chapter 6: Solution Thermodynamics and Principles of Phase Equilibria In all the preceding chapters we have focused primarily on thermodynamic systems comprising pure substances. However, in all of nature, mixtures are ubiquitous.

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